



IDEAL INSTITUTE OF PHARMACY

Pimplas Road, Vill- Posheri, Tal- Wada, Dist- Palghar

2.3.1 Student centric methods, such as experiential learning, participative learning and problem-solving methodologies are used for enhancing learning experiences

SR. NO.	PARTICULARS	Page no
1	Experiential learning	02 - 10
2	Participative learning	11 - 12
3	Problem Solving methodologies	16 - 15

PRINCIPAL

Ideal Institute Of Pharmacy
At-post-posher, Tal- Wada,
Dist - Palghar Maharashtra 421303

PRINCIPAL

Ideal Institute Of Pharmacy
At-post-posher, Tal- Wada,
Dist - Palghar Maharashtra 421303

IDEAL INSTITUTE OF PHARMA WADA

EXPERIMENTAL LEARNING

LABORATORY
JOURNAL

Subject : PIC

Name : Kedar Aman Sarda

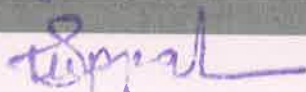
Class/Sem : F.Y.B Pharm

Roll No. : 75 Batch : B

Examination Seat No :

PRN No.

2


PRINCIPAL
Ideal Institute Of Pharmacy
At-post-posher, Tal. Wada,
Dist - Palghar Maharashtra 421303

IDEAL FOUNDATION'S
IDEAL INSTITUTE OF
PHARMACY



Certificate

This is to certify that Mr. / Miss / Mrs. Aman Sanjay
Yadav Exam seat No.

has satisfactory carried out the required practical work prescribed by the
University of Mumbai for year 2023-2024 Semester I
of B. Pharmacy course in academic year 2023-2024 and
that this journal represent his / her work done in

Place :- Posheri

Date :- 28/12/23

Internal Examiner

Principal

External Examiner



IDEAL FOUNDATION

IDEAL INSTITUTE OF PHARMACY

Instructions For Practical Work

1. Attend the laboratory punctually and regularly along with required practical kit.
2. Wear your clean apron, cap and mask before entering the laboratory. Keep hand glove with you.
3. Carefully listen, the instructions given by teacher. Make correct observations and record them immediately in the notebook.
4. Report the practical result/results on same day & get remark by teacher.
5. Work place must be clean and orderly to be safe, efficient and compatible to work. Never pick broken glasses up with your bare hands.
6. Enter the date on right corner of page.
7. Replace glassware's / laboratory reagents and chemicals on the appropriate places immediately after use with their label facing the front. After uses close the containers of chemical & reagent properly.
8. Keep your laboratory manual carefully and do not lose it as it represents your bonafide work for the year.
9. Unless your laboratory manual is certified, you are not allowed to appear for the semester practical examination.
10. You have to present the laboratory manual to the examiner for assigning marks by Him/Her.
11. Do not enter the laboratory with empty stomach. Feed yourself properly before coming for practical, otherwise the smell of chemicals will cause dizziness, vomiting, headache etc.

PRINCIPAL

Ideal Institute Of Pharmacy

At-post-poshar Tal-Wada,

Dist - Palghar Maharashtra 421303

IDEAL INSTITUTE OF PHARMACY

Course: _____

PRN No: _____

Semester: _____

Exam Seat No: _____

Sr. No.	Date	Title of Experiment	Page No.	Marks	Signature
1		Introduction to glassware and equipment			
2		Limit test for chloride			
3		Limit test for sulphate			
4		modified Limit test for chloride			
5		Limit test for iron			
6		To prepare and submit potash alum			
7		Swelling Pow Pt in Bentonite			
8		To prepare and submit boric acid			
9		To prepare and submit barium sulphate			
10		Identification test for ferrous sulphate			
11		Identification test for Copper sulphate			

5

PRINCIPAL

Ideal Institute Of Pharmacy
 At-post-postal Tal- Wada,
 Dist - Palghar Maharashtra 421303

Introduction to Glassware & Equipments

In chemistry a variety of glassware and techniques are used for the preparation, separation and purification of organic compounds.

1) Flasks:-

These are the common type of glassware used for refluxing & distillation. The different types of flasks are:

2) Condensers:-

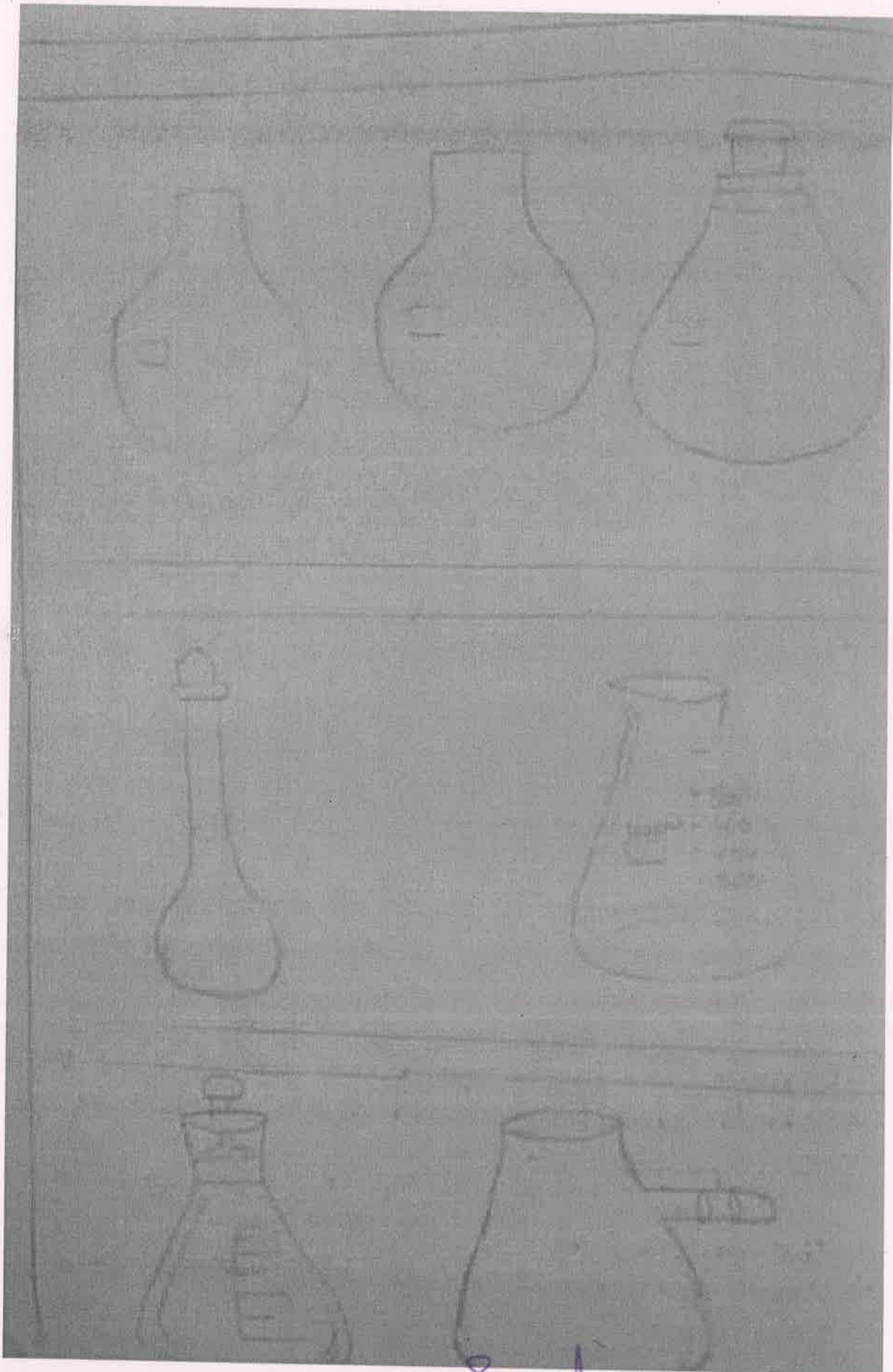
Air condenser and water condenser are used for refluxing & also ordinary distillation. The air condenser is employed if the liquid has very high boiling point.

3) Funnels: Types:-

1) Ordinary 2) Buchner 3) Separatory funnel

4) Adapters:-

These are normally used for delivery of distillate from the condenser to the receiver. Vacuum also can be applied to the adapter if required.

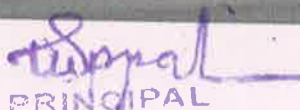


7 *Supna*
PRINCIPAL
Ideal Institute Of Pharmacy
At-post-nerher Tal-Wada,
Dist - Palghar Maharashtra 421303

Good laboratory practices

- wear coat and scened laboratory apron while working
- come well prepared by reading the principal and procedure for the experiment concerned
- Read the procedure and for consult lab in charge for solving your problem or clearing your doubts
- Do not consult your classmates for doubts which often do not give the benefit
- Maintain discipline and norms
- Use clean glass ware
- Use strong acids, alkalis and other corrosive carefully
- Do not displace the reagent from their respective places
- Do not interchange pipettes / droppers from one reagent to the other without thorough cleaning
- Use the gas whenever necessary, close the knob when not required
- Prepare your own reagent for correct results
- Submit laboratory record for correction in every practical class
- Take signature from your lecturer for all the observations
- Leave the laboratory only after cleaning your work bench

6


PRINCIPAL

Ideal Institute Of Pharmacy
At-post-pusher, Tai-Wada,
Dist - Palghar Maharashtra 421303

1] Pipette

It is a glass tube which indicates at the centre to know the volume of liquid as marked on it may be cylindrical with graduation they may be of 1ml, 2ml, 5ml, 10ml, 25ml etc.

2] Burette

It is a long graduated tube with stop cock at one end. It is made up of glass or polyvinylchloride in different volumes. It can be used to transfer or measure a desired volume of liquid. Each ml of the volume of liquid can be read on the graduated surface of the burette.

3] Measuring Cylinder

It is a tall cylinder made up of thick glass and is graduated. It is available in the capacity of 5ml, 10ml, 25ml, 50ml, 100ml, etc. It is used to measure a definite volume of a liquid.

4] Beakers

These are cylindrical, flat bottomed containers available in the capacities of 25ml, 50ml, 100ml etc. The volume of the beaker is marked on the face of the beaker.

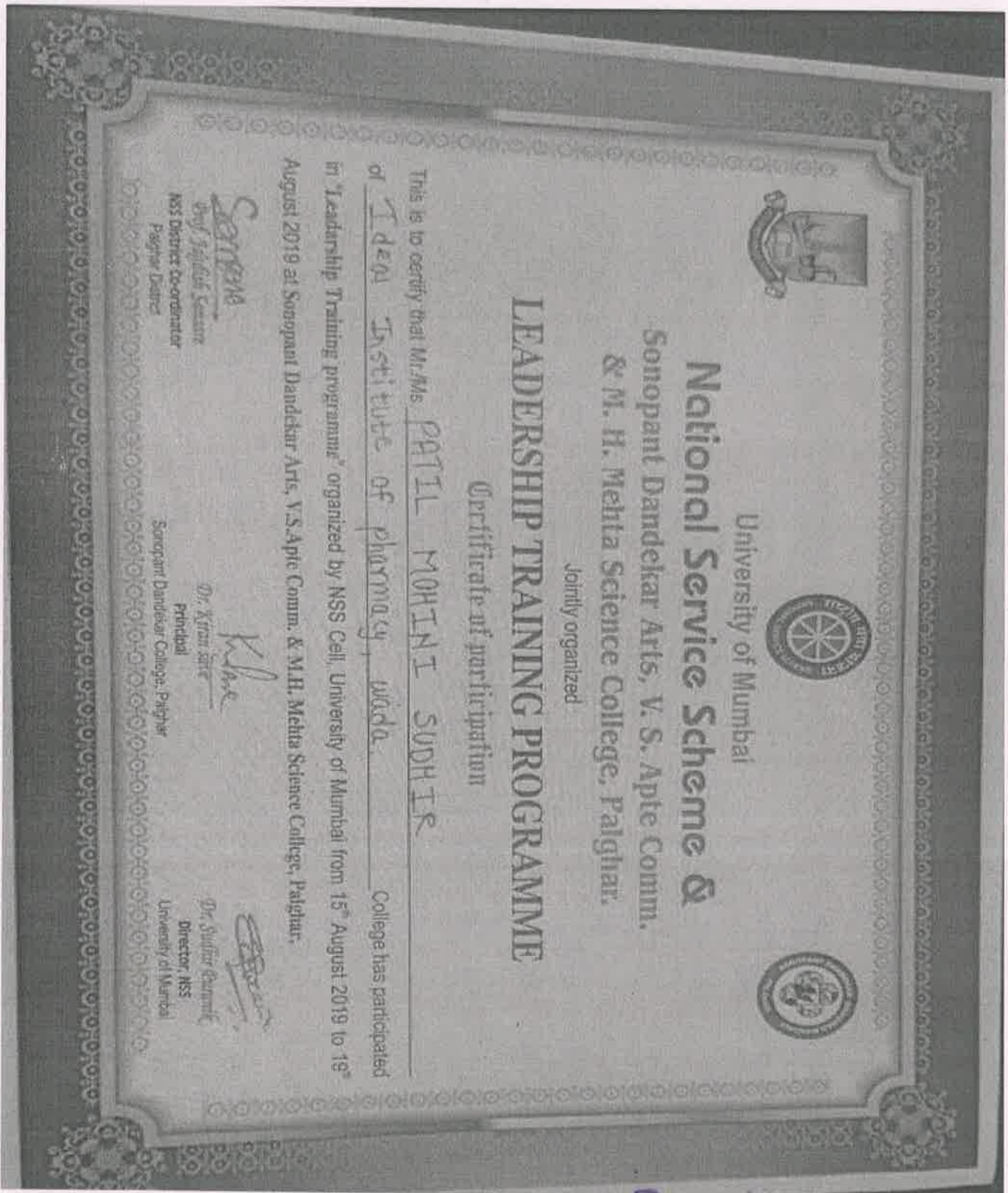

PRINCIPAL

Ideal Institute of Pharmacy
At-post-posher, Tal-Wada,
Dist - Palghar Maharashtra 421303

19	20	21	22	23	24	25	26	27	28	29	30	31	Remar	
													0	10
									P		P		33.33	11
									P		P		50.00	12
											P	P	33.33	13
					P	P			P		P		33.33	14
					P					P	P		50.00	15
													00	16
					P	P					P	P	33.33	17
													00	18
													00	19
													00	20
									P		P		33.33	21
									P		P		33.33	22
											P		33.33	23
									P	P	P	P	66.66	24

10
 PRI
 Institute of Pharmacy
 At Post-Posher, Tal. Wada,
 Dist - Paigal, Maharashtra 421303

Participative learning





University of Mumbai



National Service Scheme &

Sonopant Dandekar Arts, V. S. Aple Comm.
& M. H. Mehta Science College, Palghar.

Jointly organized

LEADERSHIP TRAINING PROGRAMME

Certificate of participation

This is to certify that Mr/Ms. Priya Vivek Murlidhar Biradari
of Ideal Institute of Pharmacy, Talada.

Leadership Training programme organized by NSS Cell, University of Mumbai from 30th August 2018 to 03rd September 2018 at Sonopant Dandekar Arts, V.S.Aple Comm. & M.H. Mehta Science College, Palghar.

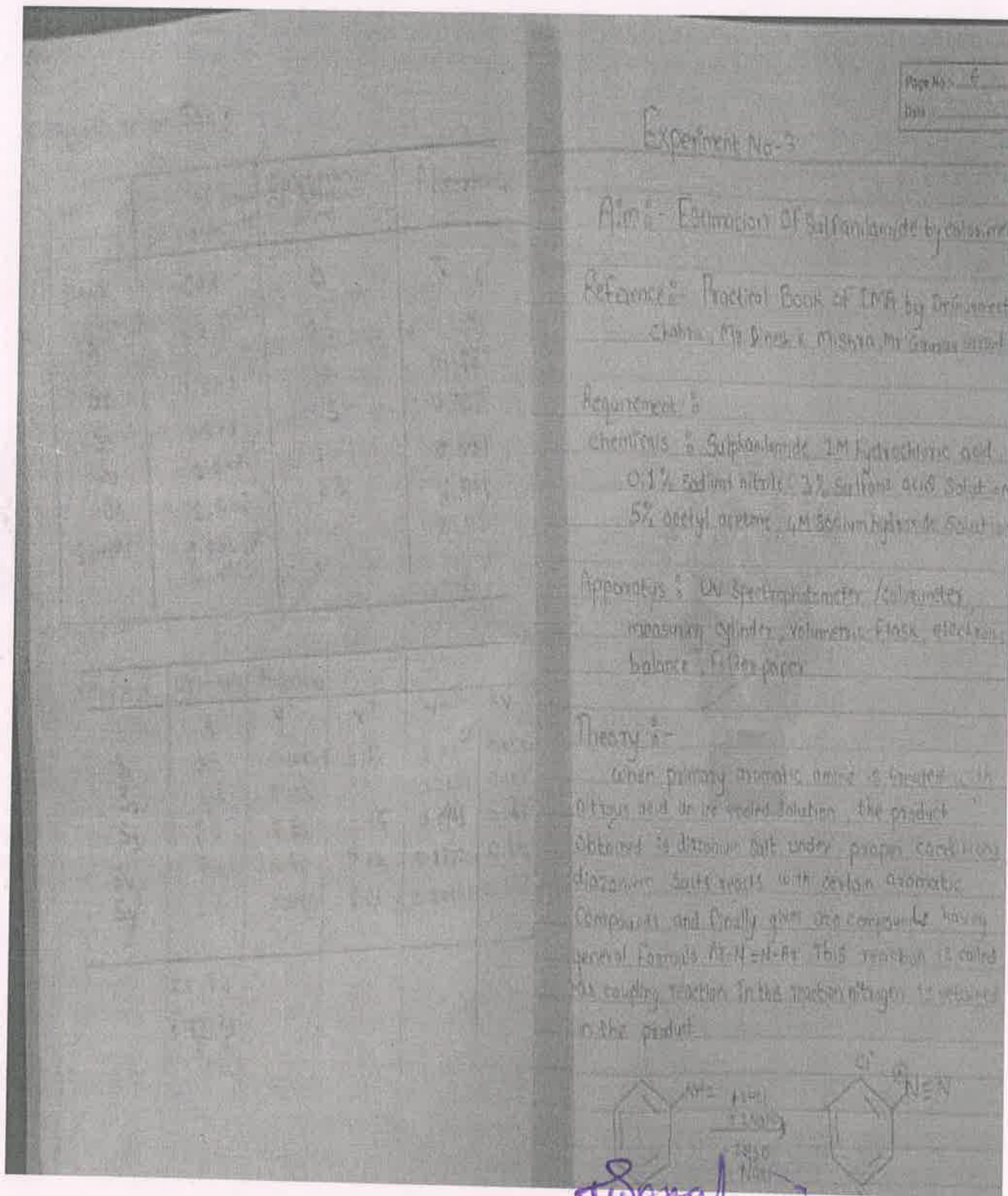
Prof. Shankar V. Yash
NSS Co-ordinator
Palghar District

Dr. Anurag Kulkarni
Principal
Sonopant Dandekar College, Palghar

Prof. V.S. Bhikse
NSS Programme Co-ordinator,
University of Mumbai

12/09/2018
Principal
Ideal Institute Of Pharmacy
At-post-poster, Tal- Wada,
Dist - Palghar Maharashtra 421303

Problem Solving methodologies



Sulphacetamide Sodium when reacted with sodium hydroxide solution in acidic medium requires hydrolysis reaction and gives sodium salt. This disodium chloride salt reacts with the active moieties groups of corticosteroids to form an ionic bond in the alkaline medium. The formed yellow color dye is the basis of quantitative estimation of sulphacetamide Sodium.

Procedure:

- 1) Sulphacetamide Sodium stock solution (1mg/ml) was prepared in distilled water.
- 2) Solutions of lower concentration were prepared by diluting the standard stock solution.
- 3) Aliquots of standard solution of sulphacetamide sodium were suitably diluted with distilled water to get concentration in range 2-100 µg/ml.
- 4) 10 ml of each dilution was transferred into series of 25ml volumetric flasks & fixed amount of water added to each of them.
- 5) Dilution was maintained below 5ml of ice water. 0.1% sodium nitrite solution was added & precipitate with continuous stirring.

1.11×10^{-5}
 $1.11 \times 10^{-5} \times 1000$
 1.11×10^{-2}
 1.11×10^{-2}
 1.11×10^{-2}
 1.11×10^{-2}
Concentration of unknown sample = 1.33

- After 3min, 1ml of 3% sulfamic acid solution was added to each flask and set aside for 5min.
- Then 4ml of 4M acetyl bromine and 4ml of 4M sodium hydroxide solution were added consecutively to each flask and mixed well.
- After 5min the analyzer was scanned by UV-visible Spectrophotometer at 420nm and the maximum absorbance was found at 420nm.
- Absorbance of all above dilution was measured against reagent at 420nm.
- Calibration curve was prepared by plotting absorbance versus concentration ($\mu\text{g}/\text{ml}$)

Result :-

The concentration of unknown sample is found to be 1.33 $\mu\text{g}/\text{ml}$